

ĐÁP ÁN ĐỀ THI HỌC KÌ 2 - NĂM HỌC 2022 - 2023
MÔN HÓA HỌC 11

I. TRẮC NGHIỆM: 7 ĐIỂM mỗi câu đúng 1/3đ

Câu	MÃ 111	MÃ 112	MÃ 113	MÃ 114
1	C	B	C	D
2	A	C	D	B
3	D	B	A	D
4	B	A	A	A
5	A	C	A	B
6	D	D	B	A
7	A	A	A	C
8	B	B	A	D
9	B	A	D	B
10	D	A	C	D
11	D	B	D	D
12	C	C	B	D
13	C	B	B	D
14	C	D	A	D
15	D	A	C	D
16	D	C	D	A
17	C	C	B	D
18	C	C	C	A
19	B	C	B	D
20	D	C	A	B
21	A	A	B	B

II. TỰ LUẬN: 3 ĐIỂM

CÂU	ĐÁP ÁN	ĐIỂM
1	<p>Viết đúng mỗi PTHH được: 0,25 điểm</p> <p>a. $\text{CH}_2=\text{CH}_2 + \text{HCl} \rightarrow \text{CH}_3\text{CH}_2\text{Cl}$</p> <p>b. $2\text{CH}_3-\text{CH}_2-\text{OH} \xrightarrow[\text{140}^\circ]{\text{H}_2\text{SO}_4 \text{ đ.}}$ $\text{CH}_3-\text{CH}_2-\text{O}-\text{CH}_2-\text{CH}_3 + \text{H}_2\text{O}$</p> <p>c. $\text{C}_6\text{H}_5\text{OH} + \text{NaOH} \longrightarrow \text{C}_6\text{H}_5\text{ONa} + \text{H}_2\text{O}$</p> <p>d. $\text{CH}_3 - \text{CH}_2 - \text{CH}=\text{O} + 2\text{AgNO}_3 + 3\text{NH}_3 + \text{H}_2\text{O} \xrightarrow{t^\circ} \text{CH}_3 - \text{CH}_2 - \text{COONH}_4 + 2\text{Ag} \downarrow + 2\text{NH}_4\text{NO}_3$</p>	<p>0,25</p> <p>0,25</p> <p>0,25</p> <p>0,25</p>
2	<p>$n_{\text{C}_2\text{H}_4} + 2n_{\text{C}_2\text{H}_2} = n_{\text{Br}_2} = 0,5$</p> <p>$\text{X} + \text{AgNO}_3/\text{NH}_3$ dư chỉ có C_2H_2 p/ư $\Rightarrow n$ kết tủa $\text{C}_2\text{Ag}_2 = n_{\text{C}_2\text{H}_2} = 0,15 \text{ mol}$</p> <p>$\Rightarrow n_{\text{C}_2\text{H}_4} = 0,2 \text{ mol} \Rightarrow m_{\text{X}} = 0,2 \cdot 28 + 0,15 \cdot 26 = 9,5 \text{ gam}$</p> <p>$\Rightarrow \% \text{C}_2\text{H}_4 = 58,95\% ; \% \text{C}_2\text{H}_2 = 41,05\%$</p>	<p>0,25</p> <p>0,25</p> <p>0,25</p> <p>0,25</p>
3	<p>$n_{\text{CH}_3\text{OH}} = 2a; n_{\text{C}_2\text{H}_5\text{OH}} = 3a$</p> <p>$\rightarrow 32.2a + 46.3a = 20,2 \rightarrow a = 0,1$</p> <p>$\rightarrow n_{\text{CH}_3\text{OH}} = 0,2 \text{ mol}; n_{\text{C}_2\text{H}_5\text{OH}} = 0,3 \text{ mol}$</p> <p>Vi H=75% nên $n_{\text{CH}_3\text{OH}_{\text{pur}}} = 0,2.75\% = 0,15 \text{ mol}$</p> <p>$n_{\text{C}_2\text{H}_5\text{OH}_{\text{pur}}} = 0,3.75\% = 0,225 \text{ mol}$</p> <p>Ta có sơ đồ $\text{CH}_3\text{OH} \rightarrow \text{HCHO} \rightarrow 4 \text{ Ag}$</p> <p>$\text{CH}_3\text{CH}_2\text{OH} \rightarrow \text{CH}_3\text{CHO} \rightarrow 2 \text{ Ag}$</p> <p>Theo sơ đồ trên suy ra $n_{\text{Ag}} = 4.0,15 + 2.0,225 = 1,05 \text{ mol}$</p> <p>$\rightarrow m_{\text{Ag}} = 113,4\text{g}$</p> <p>* Ta có $m_{\text{Y}} = m_{\text{X}} + m_{\text{O}(\text{CuO}_{\text{pur}})} = 20,2 + 16(0,15+0,225) = 26,2 \text{ gam}$</p> <p>$n_{\text{Y}} = n_{\text{Xbd}} + n_{\text{hh ancol}_{\text{pur}}} = 0,5 + 0,15+0,225 = 0,875 \text{ mol}$</p> <p>Suy ra $M_{\text{Y}} = 26,2/0,875 = 29,94 \rightarrow d_{\text{Y}/\text{H}_2} = 14,97$</p>	<p>0,25</p> <p>0,25</p> <p>0,25</p> <p>0,25</p>